Version 1.1

**General Certificate of Education (A-level)**
**June 2012**

# Chemistry CHEM5

# (Specification 2420)

# Unit 5: Energetics, Redox and Inorganic Chemistry

# Final

# *Mark Scheme*

Mark schemes are prepared by the Principal Examiner and considered, together with the relevant questions, by a panel of subject teachers. This mark scheme includes any amendments made at the standardisation events which all examiners participate in and is the scheme which was used by them in this examination. The standardisation process ensures that the mark scheme covers the candidates' responses to questions and that every examiner understands and applies it in the same correct way. As preparation for standardisation each examiner analyses a number of candidates' scripts: alternative answers not already covered by the mark scheme are discussed and legislated for. If, after the standardisation process, examiners encounter unusual answers which have not been raised they are required to refer these to the Principal Examiner.

It must be stressed that a mark scheme is a working document, in many cases further developed and expanded on the basis of candidates' reactions to a particular paper. Assumptions about future mark schemes on the basis of one year's document should be avoided; whilst the guiding principles of assessment remain constant, details will change, depending on the content of a particular examination paper.

Further copies of this Mark Scheme are available from: aqa.org.uk



Set and published by the Assessment and Qualifications Alliance.

The Assessment and Qualifications Alliance (AQA) is a company limited by guarantee registered in England and Wales (company number 3644723) and a registered charity (registered charity number 1073334).
Registered address: AQA, Devas Street, Manchester M15 6EX.

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 1(a) | To prevent it coming into contact/reacting with oxygen/air | 1 | Allow because it reacts with air/oxygen<br>And because with air/oxygen it forms an oxide. (Oxide, if identified, must be correct :- $P_4O_{10}$, $P_2O_5$, $P_4O_6$, $P_2O_6$) |
| 1(b) | One molecule contains 4P and 10O/the molecular formula is $P_4O_{10}$ | 1 | Allow exists as $P_4O_{10}$<br>Do not allow reference to combination of two $P_2O_5$ molecules<br>Ignore any reference to stability |
| 1(c) | $P_4O_{10}$ is a bigger molecule (than $SO_3$)/greater $M_r$/more electrons/ greater surface area<br>Van der Waals / vdW forces between molecules are stronger/require more energy to break | 1<br>1 | Penalise $SO_2$ for one mark (max 1)<br>CE = 0 if mention of hydrogen bonding/ionic/ giant molecule/breaking of covalent bonds<br>Do not allow just more vdW forces<br>Ignore any reference to dipole-dipole forces |
| 1(d) | $P_4O_{10} + 6H_2O \rightarrow 4H_3PO_4$<br>pH must be in the range -1 to +2 | 1<br>1 | Allow correct ionic equations<br>Ignore state symbols<br>Allow -1 to +2<br>Mark independently |

| | | | |
|---|---|---|---|
| 1(e)(i) | $3MgO + 2H_3PO_4 \rightarrow Mg_3(PO_4)_2 + 3H_2O$<br>OR $MgO + 2H_3PO_4 \rightarrow Mg(H_2PO_4)_2 + H_2O$<br>OR $MgO + H_3PO_4 \rightarrow MgHPO_4 + H_2O$ | 1 | Allow $MgO + 2H^+ \rightarrow Mg^{2+} + H_2O$<br>Allow magnesium phosphates shown as ions and ionic equations<br>Ignore state symbols |
| 1(e)(ii) | MgO is sparingly soluble/insoluble/weakly alkaline | 1 | Excess/unreacted MgO can be filtered off/separated |
| 1(e)(iii) | An excess of NaOH would make the lake alkaline/toxic/kill wildlife | 1 | Allow pH increases |

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 2(a) | $\Delta G = \Delta H - T\Delta S$ | 1 | Ignore ~~o~~ |
| 2(b) | 0.098 or 98 | 1 | Allow 0.097 to 0.099/97 to 99<br>Allow 0.1 only if 0.098 shown in working |
| | kJ $K^{-1}$ $mol^{-1}$ J $K^{-1}$ $mol^{-1}$ | 1 | Allow in any order<br>Unless slope is approx. 100(90-110) accept only kJ $K^{-1}$ $mol^{-1}$. If no slope value given, allow either units |
| | $-\Delta S/\Delta S$ | 1 | |
| 2(c) | $\Delta G$ becomes <u>negative</u> | 1 | Mark independently unless $\Delta G$ +ve then CE = 0 |
| | So reaction becomes spontaneous/feasible | 1 | Or reaction can occur below this temperature<br>Or reaction is not feasible above this temperature |
| 2(d) | Ammonia liquefies (so entropy data wrong/different) | 1 | Allow any mention of <u>change</u> in state or implied change in state even if incorrect<br>eg freezing/boiling |

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 3(a) | Enthalpy change/heat energy change when one mole of gaseous atoms<br>Form (one mole of) gaseous negative ions (with a single charge) | 1<br>1 | Allow explanation with an equation that includes state symbols<br>If ionisation/ionisation energy implied, CE=0 for both marks<br>Ignore conditions |
| 3(b) | Fluorine (atom) is smaller than chlorine/shielding is less/ outer electrons closer to nucleus<br>(Bond pair of) electrons attracted more strongly to the nucleus/protons | 1<br>1 | Fluorine molecules/ions/charge density CE=0 for both marks |
| 3(c) | Fluoride (ions) smaller (than chloride) / have larger charge density<br>So (negative charge) attracts (δ+ hydrogen on) water more strongly | 1<br>1 | Any reference to electronegativity CE=0<br>Allow H on water, do not allow O on water<br>Allow $F^-$ hydrogen bonds to water, chloride ion does not<br>Mark independently |

| | | | |
|---|---|---|---|
| 3(d)(i) | $\Delta H$(solution) = LE + $\Sigma$(hydration enthalpies) / correct cycle<br><br>*LE* = -20 -(-464 + -506)<br>= (+) 950 kJ $mol^{-1}$ | 1<br><br>1<br>1 | $AgF_2$ or other wrong formula CE = 0<br>Ignore state symbols in cycle<br><br>Ignore no units, penalise M3 for wrong units<br>-950 scores max 1 mark out of 3<br>990 loses M3 but M1 and M2 may be correct<br>808 is transfer error (AE) scores 2 marks<br>848 max 1 if M1 correct<br>1456 CE=0 (results from $AgF_2$) |
| 3(d)(ii) | There is an increase in the number of particles / more disorder / less order | 1 | Allow incorrect formulae and numbers provided number increases<br>Do not penalise reference to atoms/molecules<br>Ignore incorrect reference to liquid rather than solution |
| 3(d)(iii) | Entropy change is positive/entropy increases and enthalpy change negative/exothermic<br>So $\Delta G$ is (always) negative | 1<br><br>1 | |

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 4(a) | $\Delta H = \Sigma(\Delta H_f$ products) - $\Sigma(\Delta H_f$ reactants)<br>/= +34 - +90<br>= -56 kJ $mol^{-1}$ | 1<br><br>1 | Allow correct cycle<br><br>Ignore no units, penalise incorrect units |
| 4(b) | $\Delta S = \Sigma(S$ products) - $\Sigma(S$ reactants)<br>/= 240 - (205 +211/2)<br>= -70.5 J $K^{-1}$ $mol^{-1}$ / -0.0705 kJ $K^{-1}$ $mol^{-1}$ | 1<br><br>1 | <br><br>Ignore no units, penalise incorrect units<br>Allow -70 to -71/-.070 to -.071 |
| 4(c) | $T = \Delta H/\Delta S$ / $T$ = (Ans to part(a) $\times$1000)/ans to part(b)<br>/= -56/(-70.5 $\div$ 1000)<br>= 794 K (789 to 800 K) | 1<br><br>1 | Mark consequentially on answers to parts (a) and (b)<br>Must have correct units<br>Ignore signs; allow + or – and –ve temps |
| 4(d) | Temperatures exceed this value | 1 | |
| 4(e) | $N_2 + O_2 \rightarrow 2NO$ | 1 | Allow multiples |
| 4(f) | there is no change in the number of moles (of gases)<br><br>So entropy/disorder stays (approximately) constant / entropy/disorder change is very small / $\Delta S$=0 / $T\Delta S$=0 | 1<br><br>1 | Can only score these marks if the equation in (e) has equal number of moles on each side<br>Numbers, if stated must match equation |

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 5(a) | Electron acceptor / gains electrons / takes electrons away | 1 | Do not allow electron pair acceptor / gain of electrons / definition of redox (QWC) |
| 5(b) | $Cd(OH)_2$<br>Species (on LHS) with the least positive/most negative electrode potential / lowest *E* / smallest *E* | 1<br>1 | Do not allow '$Cd(OH)_2$/Cd'<br>Only allow this mark if M1 answer given correctly or blank<br>Do not allow negative emf |
| 5(c)(i) | 1.5 (V) / 1.50 | 1 | |
| 5(c)(ii) | $2MnO_2 + 2H_2O + Zn \rightarrow 2MnO(OH) + 2OH^- + Zn^{2+}$ | 1 | Ignore state symbols<br>$e^-$ must be cancelled<br>(take care that $Zn^{2+}$ is on RHS) |
| 5(c)(iii) | Allows ions to pass (through it) or words to that effect | 1 | Penalise passage of electrons<br>Allow mention of particular ions |
| 5(c)(iv) | Allows electrons to flow / makes electrical contact / conductor | 1 | Allow acts as an (inert) electrode / anode / cathode |
| 5(c)(v) | Zn is 'used up' / has reacted / oxidised | 1 | Allow idea that zinc reacts<br>Do not allow just zinc corrodes |

| | | | |
|---|---|---|---|
| 5(d)(i) | 3 / +3 / III<br>$2Ni(OH)_2 + Cd(OH)_2 \rightarrow 2NiO(OH) + Cd + 2H_2O$ | 1<br>1<br>1 | <br>For correct nickel and cadmium species in correct order (allow $H_2O$ missing and $OH^-$ not cancelled)<br>For balanced equation (also scores M2)<br>Allow max 1 for M2 and M3 if correct balanced equation but reversed.<br>Ignore state symbols |
| 5(d)(ii) | Metal / metal compounds are re-used / supplies are not depleted / It (the cell) can be re-used | 1 | Allow does not leak / no landfill problems / less mining / less energy to extract metals / less waste<br>Do not allow less $CO_2$ unless explained |
| 5(e)(i) | $C_2H_5OH + 3O_2 \rightarrow 2CO_2 + 3H_2O$ | 1 | Allow $C_2H_6O$ |
| 5(e)(ii) | $C_2H_5OH + 3H_2O \rightarrow 2CO_2 + 12H^+ + 12e^-$ | 1 | Allow $C_2H_6O$ |
| 5(e)(iii) | (+)0.23 (V) | 1 | |
| 5(e)(iv) | $\underline{CO_2}$ released by combustion / fermentation / fuel cell / reaction with water<br>(atmospheric) $\underline{CO_2}$ taken up in <u>photosynthesis</u> | 1<br>1 | Can be answered with the aid of equations |

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 6(a) | Co-ordinate / dative / dative covalent / dative co-ordinate | 1 | Do not allow covalent alone |
| 6(b) | (lone) pair of electrons on <u>oxygen/O</u><br>forms co-ordinate bond with <u>Fe</u> / donates electron pair to <u>Fe</u> | 1<br>1 | If co-ordination to $O^{2-}$, CE=0<br>'Pair of electrons on O donated to Fe' scores M1 and M2 |
| 6(c) | 180° / 180 / 90 | 1 | Allow any angle between 85 and 95<br>Do not allow 120 or any other incorrect angle<br>Ignore units eg $^{o}C$ |
| 6(d)(i) | 3 : 5 / 5 $FeC_2O_4$ reacts with 3 $MnO_4^-$ | 1 | Can be equation showing correct ratio |

| | | | |
|---|---|---|---|
| 6(d)(ii) | **M1** Moles of $MnO_4^-$ per titration = $22.35 \times 0.0193/1000 = \underline{4.31 \times 10^{-4}}$<br>Method marks for each of the next steps (no arithmetic error allowed for M2):<br>**M2** moles of $FeC_2O_4$= ratio from (d)(i) used correctly $\times 4.31 \times 10^{-4}$<br>**M3** moles of $FeC_2O_4$ in 250 $cm^3$ = M2 ans $\times$ 10<br>**M4** Mass of $FeC_2O_4.2H_2O$ = M3 ans $\times$ 179.8<br>**M5** % of $FeC_2O_4.2H_2O$ = (M4 ans/1.381) $\times$ 100<br>(OR for M4 max moles of $FeC_2O_4.2H_2O$ = 1.381/179.8 (= $7.68\times 10^{-3}$)<br>for M5 % of $FeC_2O_4.2H_2O$ = (M3 ans/above M4ans) $\times$ 100)<br>eg using correct ratio 5/3:<br>Moles of $FeC_2O_4 = 5/3 \times 4.31 \times 10^{-4} = 7.19 \times 10^{-4}$<br>Moles of $FeC_2O_4$ in 250 $cm^3 = 7.19 \times 10^{-4} \times 10 = 7.19 \times 10^{-3}$<br>Mass of $FeC_2O_4.2H_2O = 7.19 \times 10^{-3} \times 179.8 = 1.29$ g<br>% of $FeC_2O_4.2H_2O$ = $1.29 \times 100/1.381 = 93.4$ (allow 92.4 to 94.4)<br>Note correct answer ( 92.4 to 94.4) scores 5 marks | 1<br><br><br>1<br>1<br>1<br>1 | Allow $\underline{4.3 \times 10^{-4}}$ ( 2 sig figs)<br>Allow other ratios as follows:<br>eg from given ratio of 7/3<br><br>**M2** = $7/3 \times 4.31 \times 10^{-4} = 1.006 \times 10^{-3}$<br>**M3** = $1.006 \times 10^{-3} \times 10 = 1.006 \times 10^{-2}$<br>**M4** = $1.006 \times 10^{-2} \times 179.8 = 1.81$ g<br>**M5** = $1.81 \times 100/1.381 = 131$ % (130 to 132)<br><br>Allow consequentially on candidates ratio<br>eg **M2** = $5/2 \times 4.31 \times 10^{-4} = 1.078 \times 10^{-3}$<br>**M3** = $1.0078 \times 10^{-3} \times 10 = 1.078 \times 10^{-2}$<br>**M4** = $1.078 \times 10^{-2} \times 179.8 = 1.94$ g<br>**M5** = $1.94 \times 100/1.381 = 140$ % (139 to 141)<br>Other ratios give the following final % values<br>1:1 gives 56.1% (55.6 to 56.6)<br>5:1 gives 281% (278 to 284)<br>5:4 gives 70.2% (69.2 to 71.2) |

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 7(a) | Orange dichromate | 1 | Allow max 2 for three correct colours not identified to species but in correct order |
| | Changes to purple / green / ruby / red-violet / violet Chromium(III) | 1 | |
| | (Note green complex can be $[Cr(H_2O)_5Cl]^{2+}$ etc) | | Do not allow green with another colour |
| | That changes further to blue Chromium(II) | 1 | Allow max 1 for two correct colours not identified but in correct order |
| | $[Cr_2O_7]^{2-} + 14H^+ + 3Zn \rightarrow 2Cr^{3+} + 3Zn^{2+} + 7H_2O$ | 1 | Ignore any further reduction of $Cr^{2+}$ |
| | $2Cr^{3+} + Zn \rightarrow 2Cr^{2+} + Zn^{2+}$ / | 1 | Ignore additional steps e.g. formation of $CrO_4^{2-}$ |
| | $[Cr_2O_7]^{2-} + 14H^+ + 4Zn \rightarrow 2Cr^{2+} + 4Zn^{2+} + 7H_2O$ | | |
| 7(b) | Green precipitate | 1 | |
| | (Dissolves to form a) green solution | 1 | Solution can be implied if 'dissolves' stated |
| | $[Cr(H_2O)_6]^{3+} + 3OH^- \rightarrow Cr(H_2O)_3(OH)_3 + 3H_2O$ | 1 | Penalise $Cr(OH)_3$ once only |
| | $Cr(H_2O)_3(OH)_3 + 3OH^- \rightarrow [Cr(OH)_6]^{3-} + 3H_2O$ | 1 | Allow $[Cr(H_2O)_6]^{3+} + 6OH^- \rightarrow [Cr(OH)_6]^{3-} + 6H_2O$ |
| | | | Allow formation of $[Cr(H_2O)_2(OH)_4]^-$ and $[Cr(H_2O)(OH)_5]^{2-}$ in balanced equations |
| | | | Ignore state symbols, mark independently |

| | | | |
|---|---|---|---|
| 7(c) | (ligand) substitution / replacement / exchange | 1 | Allow nucleophilic substitution |
| | The energy levels/gaps of the <u>d</u> electrons are <u>different</u> (for each complex) | 1 | |
| | So a <u>different</u> wavelength/frequency/colour/energy of light is absorbed (when d electrons are excited) | 1 | Ignore any reference to emission of light |
| | OR light is absorbed and a different wavelength/frequency/colour/energy (of light) is transmitted/reflected | | |

| | | | |
|---|---|---|---|
| 7(d) | $E\ O_2$ (/ $H_2O$) > $E\ Cr^{3+}$ (/ $Cr^{2+}$) / e.m.f = 1.67 V | 1 | Allow E(cell) = 1.67 |
| | So $Cr^{2+}$ ions are oxidised by oxygen/air | 1 | Allow any equation of the form: $Cr^{2+} + O_2 \rightarrow Cr^{3+}$ |
| | With $[Cr(H_2O)_6]^{2+}$ get $CrCO_3$ | 1 | If named must be chromium(II) carbonate |
| | with $[Cr(H_2O)_6]^{3+}$ get $Cr(H_2O)_3(OH)_3$ / $Cr(OH)_3$ | 1 | Allow 0 to 3 waters in the complex |
| | and $CO_2$ | 1 | Can score M3, M4, M5 in equations even if unbalanced |
| | Cr(III) differs from Cr(II) because it is acidic / forms $H^+$ ions | 1 | |
| | because $Cr^{3+}$ ion polarises <u>water</u> | 1 | Ignore charge/size ratio and mass/charge |

| Question | Marking Guidance | Mark | Comments |
|---|---|---|---|
| 8(a) | | | For reactions 1 to 3 must show complex ions as reactants and products<br>Take care to look for possible identification on flow chart |
| | **Reaction 1** | | |
| | ammonia solution | 1 | |
| | **W** is $[Co(NH_3)_6]^{2+}$ | 1 | |
| | $[Co(H_2O)_6]^{2+} + 6NH_3 \rightarrow [Co(NH_3)_6]^{2+} + 6H_2O$ | 1 | Correct equation scores all 3 marks |
| | **Reaction 2** | | Allow oxygen, Do not allow air |
| | $H_2O_2$ | 1 | |
| | **X** is $[Co(NH_3)_6]^{3+}$ | 1 | |
| | $2[Co(NH_3)_6]^{2+} + H_2O_2 \rightarrow 2[Co(NH_3)_6]^{3+} + 2OH^-$ | 1 | Allow $2[Co(NH_3)_6]^{2+} + {}^1/_2O_2 + H_2O \rightarrow 2[Co(NH_3)_6]^{3+} + 2OH^-$<br>Correct equations score all 3 marks |
| | **Reaction 3** | | |
| | HCl | 1 | Do not allow $Cl^-$ but mark on |
| | **Y** is $[CoCl_4]^{2-}$ | 1 | |
| | $[Co(H_2O)_6]^{2+} + 4Cl^- \rightarrow [CoCl_4]^{2-} + 6H_2O$/ | 1 | Correct equation scores previous mark |
| | $[Co(H_2O)_6]^{2+} + 4HCl \rightarrow [CoCl_4]^{2-} + 6H_2O + 4H^+$ | | This equation scores all three marks |

| | | | |
|---|---|---|---|
| | **Reaction 4**<br>$Na_2CO_3$ Or $NaOH/NH_3$<br>**Z** is $CoCO_3$ $Co(OH)_2/Co(H_2O)_4(OH)_2$<br>$[Co(H_2O)_6]^{2+} + CO_3^{2-} \rightarrow CoCO_3 + 6H_2O$ $[Co(H_2O)_6]^{2+}+2OH^- \rightarrow Co(H_2O)_4(OH)_2+2H_2O$ etc<br>Or $[Co(H_2O)_6]^{2+} + Na_2CO_3 \rightarrow CoCO_3 + 6H_2O + 2Na^+$ | 1<br>1<br>1 | Do not allow $CaCO_3$ as a reagent but mark on<br>Allow waters to stay co-ordinated to Co. This mark also previous mark<br>Allow $Co^{2+} + CO_3^{2-} \rightarrow CoCO_3$ |
| 8(b) | $SO_3^{2-} + {}^1/_2O_2 \rightarrow SO_4^{2-}$<br>The activation energy is lower (for the catalysed route)<br>${}^1/_2O_2 + 2Co^{2+} + 2H^+ \rightarrow H_2O + 2Co^{3+}$<br>$2Co^{3+} + SO_3^{2-} + H_2O \rightarrow 2Co^{2+} + SO_4^{2-} + 2H^+$ | 1<br>1<br>1<br>1 | Allow multiples<br>Or $Co^{3+}$ attracts $SO_3^{2-}$/$Co^{2+}$ attracts $SO_3^{2-}$ /oppositely charged ions attract<br>Allow these equations in either order |